

Answers

Question 1

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|---------------------------------|-----------------------|-----------------------|
| A. lithium oxide | B. potassium chromate | C. cesium chloride |
| D. Rbiodium acetate | E. Magnesium sulfide | F. Magnesium nitride |
| G. Calcium bromide tetrahydrate | H. Calcium phosphide | I. Barium iodide |
| J. Aluminum sulfide | K. potassium hydride | L. silver thiocyanate |
| M. Sodium sulfite | N. sodium chlorate | |

Question 2

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|------------------------------------|---------------------------|------------------------------|
| A. Iron (II) chloride | B. Chromium (III) oxide | C. Cobalt (III) permanganate |
| D. Cobalt (II) sulfide | E. Manganese (IV) sulfate | F. Manganese (II) dichromate |
| G. Nickel (II) bromide hexahydrate | H. Nickel (II) sulfide | I. Mercury (II) oxide |
| J. Mercury (I) oxide | K. Copper (II) nitrate | L. Manganese (V) phosphate |
| M. Tin (IV) chloride | N. Copper (I) iodide | |

Question 3

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|---|---|--------------------------------|
| A. $\text{AlN} \cdot 5 \text{H}_2\text{O}$ | B. CaC_2O_4 | C. AgMnO_4 |
| D. $\text{Na}_2\text{Cr}_2\text{O}_7$ | E. $\text{Zn}(\text{HSO}_4)_2$ | F. LiH |
| G. CdI_2 | H. $\text{Hg}_2(\text{H}_2\text{PO}_4)_2$ | I. Ba_3N_2 |
| J. $\text{ZnCl}_2 \cdot 3 \text{H}_2\text{O}$ | K. LiClO | L. CsCN |
| M. FeP | N. SnBr_2 | O. Hg_2S |
| P. Cr_2S_3 | Q. PbO | R. $\text{Au}(\text{NO}_3)_3$ |
| S. Au_3PO_4 | T. $\text{NiNO}_2 \cdot 2 \text{H}_2\text{O}$ | U. $\text{Sn}(\text{CrO}_4)_2$ |
| V. FeF_2 | W. FeI_3 | X. $\text{Pb}(\text{OH})_4$ |

Question 4

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|-------------------------|------------------------------|-------------------------|
| A. Carbon disulfide | B. Phosphorous trichloride | C. sulfur dioxide |
| D. Carbon monoxide | E. phosphorous pentachloride | F. nitrogen trichloride |
| G. Carbon tetrachloride | H. Silicon dioxide | I. Boron trichloride |
| J. Sulfur tetrafluoride | K. dinitrogen pentaoxide | L. Hydrogen bromide |
| M. Sulfur hexafluoride | D. Carbon tetrabromide | |

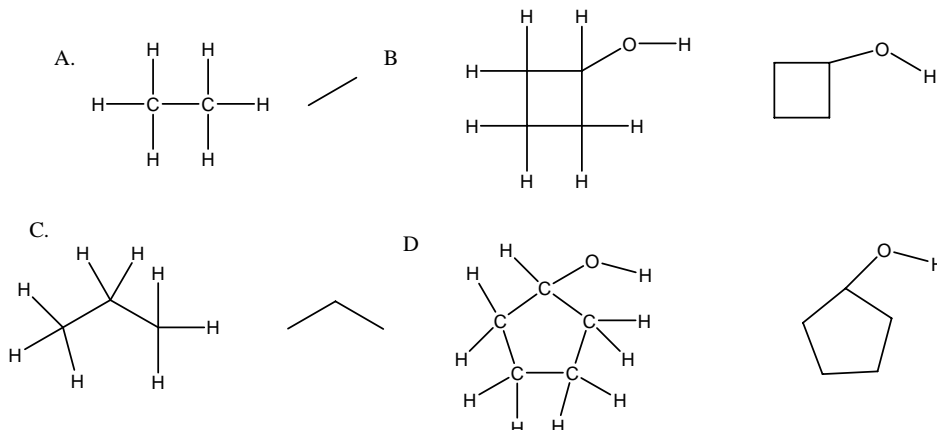
Question 5

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|------------|------------|-----------------|
| A. butanol | B. Decane | C. benzene |
| D. heptane | E. pentane | F. cyclohexanol |

Question 6

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|------------------------------|-----------------------------------|---|
| A. CH_4 | B. $\text{C}_3\text{H}_8\text{O}$ | C. C_5H_{10} |
| D. C_8H_{18} | E. C_6H_6 | F. $\text{C}_6\text{H}_{11}\text{ClOS}$ |

Question 7

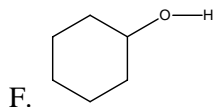
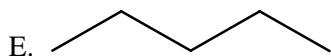


5. Give the molecular formula for the following molecular compounds

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|-----------------------------------|--------------------------------------|
| A. sulfur trioxide | B. phosphorous trihydride (phosgene) |
| C. ammonium (nitrogen trihydride) | D. Oxygen difluoride |
| E. dinitrogen pentasulfide | F. tetrasulfur dinitride |
| G. nitrogen dioxide | H. Water (dihydrogen oxide) |
| I. tetraphosphorous decaoxide | J. tetra phosphorous hexaoxide |
| K. Oxygen dichloride | L. Peroxide (dihydrogen dioxide) |

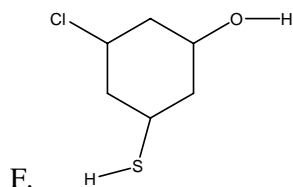
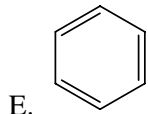
5. Give the name for the following organic compounds.

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|-----------------|-------------------|
| A. $C_4H_{10}O$ | B. $C_{10}H_{22}$ |
| C. C_6H_6 | D. C_7H_{16} |



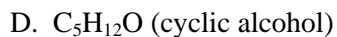
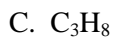
6. Write the molecular formula for the following organic compounds

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|-----------------|-------------|
| A. methane | B. propanol |
| C. cyclopentane | D. octane |



7. Draw the structural and line formulas for the following organic compounds.

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|-----------|-----------------|
| A. ethane | B. cyclobutanol |
|-----------|-----------------|



1. Give the name for the following ionic compounds containing **fixed** charge metals:

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|----|--|----|----------------------------|
| A. | Li_2O | B. | K_2CrO_4 |
| C. | CsCl | D. | RbCH_3CO_2 |
| E. | MgS | F. | Mg_3N_2 |
| G. | $\text{CaBr}_2 \cdot 4 \text{H}_2\text{O}$ | H. | Ca_3P_2 |
| I. | BaI_2 | J. | Al_2S_3 |
| K. | KH | L. | AgSCN |
| M. | $\text{Na}_2\text{S}_2\text{O}_3$ | N. | NaClO_3 |

2. Name the following binary ionic compounds containing **variable charge** metals:

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|----|--|----|------------------------------|
| A. | FeCl_2 | B. | Cr_2O_3 |
| C. | $\text{Co}(\text{MnO}_4)_3$ | D. | CoS |
| E. | $\text{Mn}(\text{SO}_4)_2$ | F. | MnCr_2O_7 |
| G. | $\text{NiBr}_2 \cdot 6 \text{H}_2\text{O}$ | H. | NiS |
| I. | HgO | J. | Hg_2O |
| K. | $\text{Cu}(\text{NO}_3)_2$ | L. | $\text{Mn}_3(\text{PO}_4)_5$ |
| M. | SnCl_4 | N. | CuI |

3. Give the molecular formula for the following ionic compounds (note both fixed and variable charged metals)

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|----|--------------------------------|----|-----------------------------------|
| A. | Aluminum nitride penta hydrate | B. | Calcium oxylate |
| C. | Silver permanganate | D. | Sodium dichromate |
| E. | Zinc bisulfate | F. | Lithium hydride |
| G. | Cadmium iodide | H. | Mercury (II) dihydrogen phosphate |
| I. | Barium nitride | J. | Zinc chloride trihydrate |
| K. | Lithium hypochlorite | L. | Cesium Cyanide |
| M. | Iron (III) phosphide | N. | Tin (II) bromide |
| O. | Mercury (I) sulfide | P. | Chromium (III) Sulfide |
| Q. | Lead (II) oxide | R. | Gold (III) Nitrate |
| S. | Gold (I) phosphate | T. | Nickel (I) Nitrite dihydrate |
| U. | Tin (IV) chromate | V. | Iron (II) Fluoride |
| W. | Iron (III) iodide | X. | Lead (IV) hydroxide |

4. Name the following binary **molecular** compounds:

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|----|------------------------|----|----------------|
| A. | CS_2 | B. | PCl_3 |
| C. | SO_2 | D. | CO |
| E. | PCl_5 | F. | NCl_3 |
| G. | CCl_4 | H. | SiO_2 |
| I. | BCl_3 | J. | SF_4 |
| K. | N_2O_5 | L. | HBr |
| M. | SF_6 | N. | CBr_4 |